

Solutions

- Concentration & Dilution
- Micropipetting & Assays

A Solution

- Substances can react when they encounter each other
 - Must be fluid or dissolved in a fluid
- **Solution: solutes** dissolved in **solvent**
 - **Aqueous solution:** solvent is **water**
- **Concentration:** amount of a solute in the total volume of solution

Concentration

- = Number of solutes in a given volume
- Measurable units
 - Milliliters per 100ml (v%)
 - Grams per 100ml (g%)
 - Micrograms per milliliter ($\mu\text{g/ml}$)
 - Parts per thousand (ppt)
 - But a gram of sucrose has fewer molecules than a gram of glucose,
 - so 1% sucrose \neq 1% glucose concentration

Concentration

- = Number of solutes in a given volume
- “Real” units:
 - number of solute molecules per volume**
 - **Molar (M)**
 - = **Moles per liter** = millimoles per milliliter
- 1 Mole = 6.023×10^{23}
- 1M glucose = 6.023×10^{23} glucose molecules per 1 liter of total solution
 - 1 M glucose = 1M sucrose in concentration
 - = $1M = 10^3 \text{ mM} = 10^6 \mu\text{M}$

Concentration

But how can you make a “real” concentration?
You cannot count the molecules!

Conversion factor

= **molecular mass** [molecular weight (MW)]

= daltons per molecule

= **grams per mole of molecules**

e.g., $MW_{\text{glucose}} = 180$.

i.e., 1 mole of glucose weighs 180 grams

Concentration

So, say you want 10 ml of 2M glucose solution:

Remember, $MW_{\text{glucose}} = 180 \text{ g/mole}$.

$10\text{ml} \times 2 \text{ moles glucose}/1000 \text{ ml} \times 180\text{g glucose/mole glucose}$
= 3.6 g glucose needed.

What is the “real” concentration of a 1% glucose solution?

$1 \text{ g glucose}/100 \text{ ml} \times 1 \text{ mole}/180 \text{ g} \times 1000\text{ml/L}$
= $0.056 \text{ M} = 56 \text{ mM}$ glucose solution.

Solutions & Dilutions

Changing Concentration: Dilution

$$C_1 V_1 = C_2 V_2$$

C_1 : concentration of the initial solution
 V_1 : volume of the initial solution
 C_2 : concentration of the final solution
 V_2 : volume of the final solution

Changing Concentration: Dilution

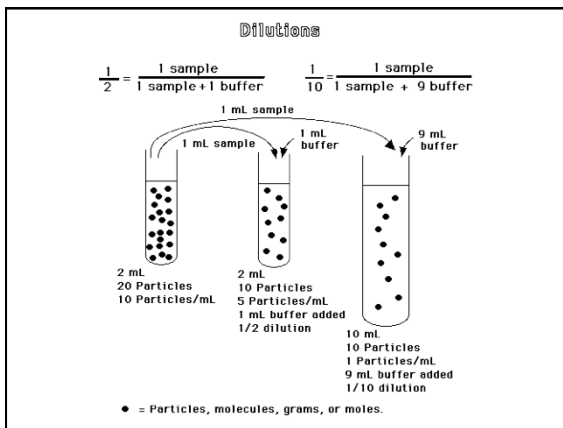
$$C_1 V_1 = C_2 V_2$$

$$\therefore C_1 = C_2 V_2 / V_1$$

$$\therefore C_2 = C_1 V_1 / V_2$$

$$\therefore V_1 = V_2 C_2 / C_1$$

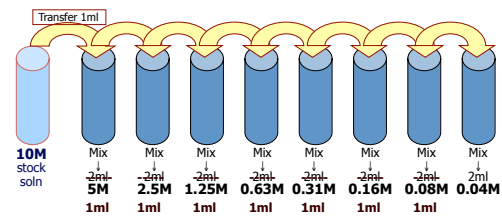
$$\therefore V_2 = V_1 C_1 / C_2$$



Changing Concentration: Serial Dilutions

To create a precise series of dilutions, or achieve a low concentration in a small volume.

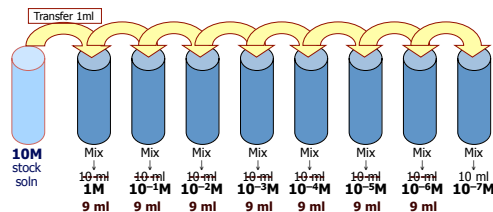
E.g., a **1:2 dilution series**: Starting stock solution = 10M.
 Place 1 ml solvent in each dilution tube



Changing Concentration: Serial Dilutions

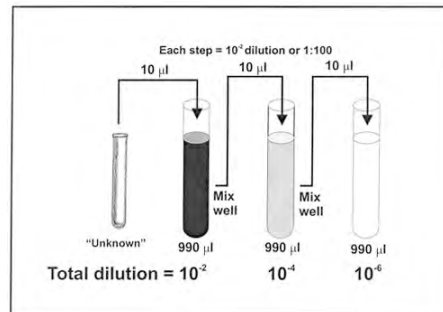
To create a precise series of dilutions, or achieve a low concentration in a small volume.

E.g., a **1:10 dilution series**: Starting stock solution = 10M.
 Place 9 ml solvent in each dilution tube



Changing Concentration: Serial Dilutions

• **1:100 dilution series**:



Solutions & Dilutions

